Absorption of DCM Dye in Ethanol: Experimental and Time Dependent Density Functional Study

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ABSTRACT— **Experimental and theoretical absorption spectra of [2-[2-[4-(dimethylamino) phenyl]ethenyl]-6-methyl-4H- pyran-4-ylidene]-propanedinitrile (DCM) have been studied. UV-Visible (UV-Vis.) absorption spectrum of DCM has been reported after its synthesis. Two relatively intense peaks appeared at 473 and 362 nm respectively. A theoretical investigation on the electronic structure of DCM is presented in an effort to rationalize our experimental results. Theoretical results have been obtained with a polarizable continuum model time-dependent density functional theory (PCM-TD-DFT) approach. At first, a vast functional benchmark has been performed to determine a suitable approach for determination of electronic structure and UV-Vis. absorption spectrum of DCM. In a second step, we evaluated the impact of the atomic basis set on the electronic transition energies using a large panel of Pople's basis sets ,up to the 6-31+G(3df,2p) and also a correlation consistent basis set, cc-pVTZ. It turns out that the selected basis set has a relatively finite influence on the calculated electronic transition energies as well as the topology of the absorption shape, but both are significantly affected by the chosen functional. In the present case, no single functional simultaneously provides highly accurate positions and intensities of the different bands, but mPW1PBE and mPW1LYP appear to be a good compromise. The mPW1PBE along with medium basis sets produced both absorption bands with maximum peaks about 463 and 346 nm. At all stages, ethanol has been chosen as a solvent environment. To improve the accuracy of first electronic excitation, a complete analysis of the origin of the band shape using TD-DFT vibrational couplings was performed. Finally the computed transition energy was corrected to 472 nm which was in excellent agreement with experiments.**

KEYWORDS: TD-DFT, DCM dye, Absorption spectra, Vibrational coupling.

I. INTRODUCTION

There has been a great deal of research on DCM ([2-[2-[4-(dimethylamino) phenyl]ethenyl]-6-methyl-4H- pyran-4-ylidene]-propanedinitrile), for its advantage of widely tunable, efficient laser dye [1]–[4] and capability to use as an active element in luminescent solar concentrators [5], [6]. The photophysical properties, such as the absorption and emission spectra provide a direct way for studying the details of electronic nature on the primary photoprocesses that are operative in a fluorescent dye. DCM has also been used as a guest in polyimide hosts for electro-optic devices [7]. The presence of an electron-donating group (e. the dimethyl amino) at 4-position in phenyl ring decreases HOMO-LUMO gap and increases the fluorescence quantum yield which has led it to its vast applications in dye lasers. Hence, the photophysics and photochemistry investigations of this dye have been studied intensively [6], [8]–[11]. Also nowadays, DCM and its analogs are considered in color graphical displays [12], inorganic light diodes [13], and as a fluorescent probe in biological systems [14].

DCM is an example of a donor-π-bridge-acceptor molecule (Fig. 1). It consists of an

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electron-donating dimethylamino group and an electron-accepting dicyano group, separated by a conjugated bridge (Fig. 1) [15]. The bridge is composed of two aromatic rings and alternating single and double bonds that extend all the way from the donor to the acceptor. Conjugation gives rise to a π molecular orbital that extends from the donor to the acceptor by overlapping of the p-orbitals from the constituent carbon atoms or from lone pairs. The amino nitrogen in the donor group of DCM is expected to have orbitals that are between pure sp^3 and sp^2 hybrid orbitals. This is due to resonance structures that involve delocalization of the one pair to the aromatic ring [16]. Cis and Trans isomers of DCM have been identified in solutions via measurement of the absolute absorption cross sections by J. M. Darke *et al.* [17]. The effect of the solvent polarity on the absorption spectra has been reported by Meyer *et al.* [9]. The evolution of DCM during electronic excitation has been investigated in both theoretical and experimental field by numerous authors [8], [18]–[26].

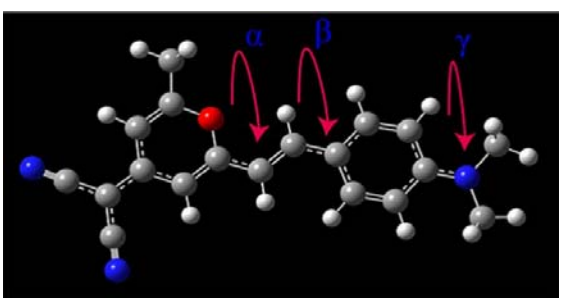

Fig. 1. [2-[2-[4-(dimethylamino)phenyl]ethenyl]-6-methyl-4H-pyran-4-ylidene]- propanedinitrile (DCM). α, β and γ are the three rotors (dihedral angles) was used to investigate for potential energy curves (PECs) on ground (S_0) and first excited (S_1) states. This picture has been depicted by gaussview 05 software.

In this article, we synthesized DCM and showed different absorption spectra in solvent environment (ethanol) in the UV-Vis. region. Two relatively intense peaks appeared at 473 and 362 nm respectively. As far as we know, there is not any theoretical investigation about effect of different DFT methods on accurate simulation of absorption spectra of solvated DCM. Our aim is to simulate the absorption spectra of DCM accurately. In order to obtain this aim, one needs to include electron correlation. In practice, this can be done with TDDFT. An advantage of TD-DFT, compared to more refined electron correlated wave function theories, is the possibility to include bulk solvation effects during excited-state force minimizations, through the use of dielectric approximations, typically the polarizable continuum model (PCM) [27], [28]. The basis sets and functionals effects for simulation of DCM spectra are investigated. Among the examined functional/ basis levels, mPW1PBE simultaneously provides accurate positions and intensities of the different bands. To improve the accuracy of simulations, we have to evaluate the vibronic couplings. Vibrational modes that hook up with the electronic transition are not clear in shape of experimental spectra but change the final position of maximum peaks in spectra. To perform such task, one needs to calculate the Hessian of both the ground state and relevant (generally first) singlet excited state. In practice, this can be done with DFT and TDDFT for ground and excited states respectively. The first TDDFT derivatives (gradients) allows to minimize the excited-state geometries since the seminal work of Van Caillie and Amos [29] and the subsequent extensions has been improved [27], [30], [31]. Therefore computed transition energies corrected to 472 nm by considering the vibrational transition during first electronic excitation. This paper is organized as follows: In Sections 2.1 and 2.2, our experimental and computational protocols are described. In Sections 3.1 and 3.2, the functionals and basis set effects for DCM absorption spectra are investigated. Section 3.3 gives the computed vertical spectra with vibrational couplings of DCM and eventually, we conclude.

II. Synthesize Method

A. *Experiment*

The chemicals for synthesizing of DCM dye were obtained from Merck and Aldrich and used without further purification. Melting

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point was analyzed with an electro thermal melting point apparatus. The hydrogen NMR (nuclear magnetic resonance) spectrum was obtained on FT-NMR (Fourier transform-NMR-500MHz) Brucker spectrometer and the chemical shifts are expressed in δ ppm using tetra methyl silane (TMS) as an internal standard.

For synthesis of 4- dicyano methylene- 2,6-dimethyl-4H- pyran (DCM intermediate), a mixture of malonitirile (1 mmol), 2,6-dimethyl-4H-pyran-4-one (1 mmol) and acetic anhydride (0.5 ml) was refluxed for 1.5 h. The unreacted acetic acid was aspirated off and the residue was washed with 50 ml. of boiling water and collected to give soft brown material. Recrystallization from boiling heptane produced a brown powder. After that, for synthesis of 4-(Dicyanomethylene)-2-methyl-6-(4-dimethylaminostyryl)-4H-pyran (DCM), the mixture of (1.06 g) (6.16 mmole) synthesized 4-dicyanomethylene-2,6-dimethyl-4H- pyran, (0.76 g) (5.048 mmol) (N,N-Dimethyl) benzaldehyde, (87 ml) toluene, 60 drops of piperidine and 90 drops of acetic acid was added. The mixture was refluxed (120° C) under argon gas for 20 h. Afterwards the precipitated dye was filtered. The product was purified by column chromatography using n-hexane- ethyl acetate (6:1) as eluent. Fig. 2 show the schematic for preparing the DCM dye. Hydrogen NMR spectrum of final product (DCM) is available in supporting information.

UV-Vis. spectra of DCM diluted in ethanol has been taken by a spectrometer (Avantes Avaspec-2048) supplied with a white light source with 200-1200 nm wavelength range (Avantes Avalight-Hall). The concentration of synthesized DCM dissolved in ethanol was 10^{-5} M/lit and the diluted DCM hold in a 1×1×5 cm^3 quartz cuvette for measuring its absorption spectrum in ambient temperature. The Figure 3 shows the experimental absorption spectra. As can be seen, the room temperature absorption spectrum of DCM exhibit two main bands in polar solvent (ethanol).

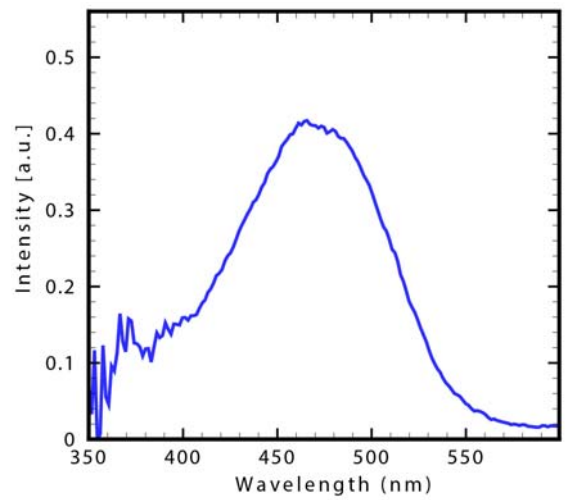

Fig. 2. Experimental UV-Vis. absorption spectrum of solvated DCM at room temperature.

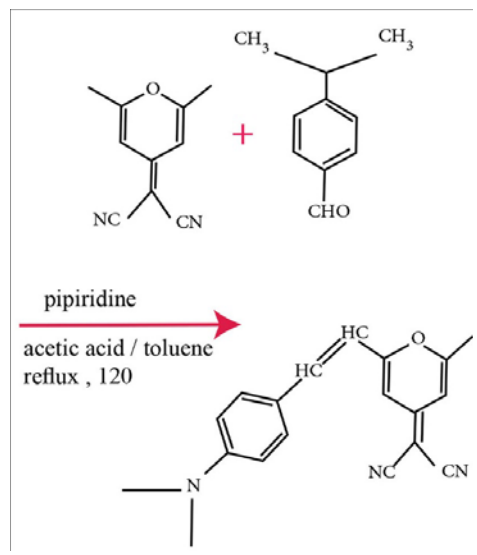

Fig. 3. DCM synthesis protocol

A low energy band spanned in the 450-520 nm region and second band that is less intense spanned in higher energy 340 -380 nm. Maximum peak in each one aforementioned bands appear in 473 and 362 nm respectively (Fig. 3).

B. *Computational Details*

All calculations have been performed with the Gaussian09 [32] program suite, using default thresholds and algorithms except when it has noted. We have selected a large set of atomic (ABS) and DFT functionals: B3LYP [33]–[35], X3LYP [36], O3LYP [37], BLYP [34]–[38], MPW1PBE [39], [40], MPW3PBE [39], [40], MPW1LYP [34]–[38], [41], [42] , M06-2X [42], B97-2 [42], wB97XD [43] and cam-

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B3LYP [44] to perform benchmarks for DCM. We have used the popular Pople's ABS [45], [46] in this paper. Triple zeta correlation consistent ABS (cc-pVTZ) [47] has been used as well. All Pople's ABS used in this work are split valence basis sets of contracted gaussian type functions (CGTFs) with different diffuse and polarization functions. The cc-pVTZ is a developed CGTF which designed by Dunning and co-workers for use in calculation methods (such as CI and DFT) that include electron correlation [47]. The ground-state, S0 (and first excited-state, S1) geometries have been optimized with DFT (TD-DFT) until the residual mean square forces were smaller than 3×10^{-4} (au). The vibrational frequencies have been computed analytically for the S0 (S1) state. They confirmed the presence of minima characterized by the absence of imaginary modes (and transition states characterized by a single imaginary mode). During all calculation steps, we have included bulk solvent effects by using the PCM [28] that correctly models the major solvent effects as long as no specific solute-solvent interactions are implied. In PCM model, the solvent is indicated by a homogeneous dielectric polarized by the solute, located within a cavity constructed as an envelope of spheres centered on solute atoms. We have selected ethanol as a polar solvent. Vibrationally resolved spectra within the harmonic approximation were computed using the FC classes program [48]–[50]. The reported spectra have been simulated at 298.15 (K) using a convoluting gaussian functions presenting a full width at half maximum (FWHM) of 0.33 (eV) for vertical excitation without vibrational corrections. FWHM for simulating of vibronic spectra was adjusted in accordance with the experimental information (about 0.33 (eV)) and fine structure of vibronic transitions (135 and 1 (cm^{-1})). A maximum quantum number that has to be considered for each vibrational mode (overtone) involved in the combination bands were 20 and 13 respectively. A maximum number of integrals to be computed for each class were set to 10^8 for comparable discussions. Experimental and all computational spectra are reported as normalized absorption line shapes.

III. Results and Discussion

Given the planarity of the aromatic rings and ylidene and the symmetry of the methyl groups, the streochemical configuration of DCM is significantly specified by the α, β and γ dihedral angles, which refer to the torsion around of two C-C and one C-N bonds respectively (Fig. 1). Ground state potential energy curves (PECs) has been obtained by variation of each torsion angle separately and calculation of electronic energies (Fig. 4). As can be seen in Fig. 5, the points A, B, D, E and F have same structure and energy. So there are two different minima in PEC, A and C that are correspond to *cis* (α≈0°) and *trans* (α ≈ 180°) isomers respectively and $\beta=\gamma=0$ as well.

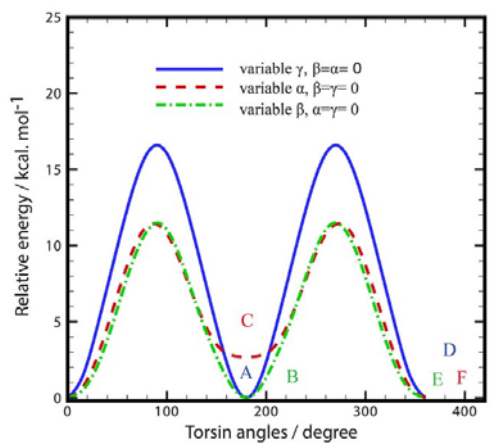

Fig. 4. Ground state potential energy curves (PECs) that has been obtained by variation of each torsion angle and electronic energy calculations, separately. All calculations have been performed at PCM-B3LYP/6-31G(d) level. Relative energy unit is ($kcal.mol^{-1}$) and angle unit is degree. A, B, D, E and F isomers have been labeled *cis* and C isomer has been labeled *trans*.

Frequency calculations on optimized *cis* and *trans* structure proved that these are in minimum energy. Relative computed thermodynamic parameters and electric dipole moments are shown in Table 1.

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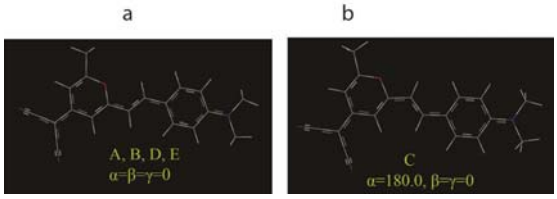

Fig. 5. (a), *cis* isomer and (b), *trans* isomer. Geometry minimizations have been performed at PCM-B3LYP/6-31G(d) level.

An energetic comparison of two minima indicates that, in the polar solvent, *cis* isomer is most favored conformer on the ground electronic state. Thus, at the following sections we focused on electronic structure and photochemical properties of this minimum.

Table 1. Computed thermodynamic parameters and dipole moments μ obtained at B3LYP/6-31G(d)// B3LYP/6-31G(d) level for *cis* and *trans*. All energetic parameters reported in (kcal.mol^{-1}) unit. Dipole moment's unit μ is Debye. Angle's unit is degree. ΔE_e and ΔE_0 are electronic energy difference without and with considering the zero-point vibrational energy respectively. ΔE^{298}, ΔH^{298} and ΔG^{298} are electronic energy, enthalpy and Gibbs free energy differences at room temperature.

Isomer	Cis	Trans
α	0.0	180
β	0.0	0.0
γ	0.0	0.0
ΔE_e	0.0	2.9
ΔE_0	0.0	2.5
ΔE^{298}	0.0	1.6
ΔH^{298}	0.0	1.6
ΔG^{298}	0.0	3.8
μ	21.6	20.4

The electronic structure of most favored DCM conformer during computation of excitations at PCM-TD-B3LYP/6-31G(d) level is considered before examination of different functionals and basis sets. Computed spectrum has been shown in Fig. 6. By comparing computed spectrum with experimental spectrum (Fig. 3) it is seen that the two aforementioned maximum peaks are resulted from singlet HOMO→LUMO and HOMO-1 →LUMO in combination with HOMO→ LUMO+1 electronic transition (Fig. 6).

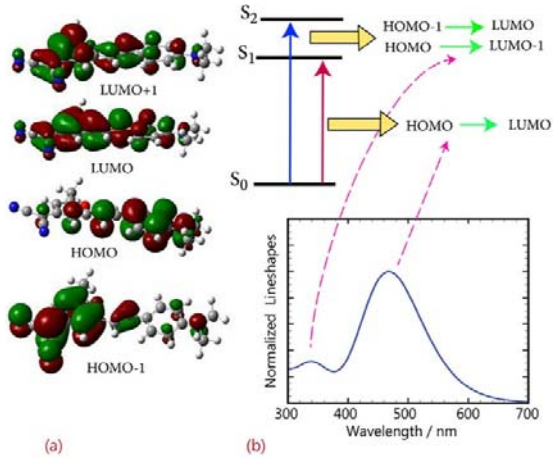

Fig. 6. (a) Canonical Orbitals which is included in first two singlet electronic excitation processes. (b), Computed absorption spectrum of solvated DCM resulted from vertical excitation computation without vibrational couplings. Schematic diagram depicts each of transition processes. All computation have been performed at PCM-TD-B3LYP/6-31G(d) level.

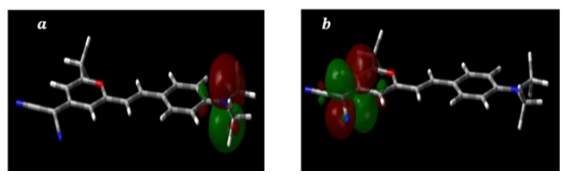

Fig. 7. Localized natural bond orbitals: a) Nature of HOMO is $\pi_{c-N_{amino}}$ b) Nature of LUMO is π^*_{c-c} .

As can be seen in Fig. 6, nature of desired orbitals is π and π*. Natural bond orbital analysis shows that the HOMO and LUMO have $\pi_{c-N_{amino}}$ bonding and π^*_{c-c} antibonding nature (Fig. 7). It means that intermolecular charge transfer occurs during first electronic excitation. This phenomenon has been proposed for photo-excitation of D-π-conjugated bridge-A systems by many research groups [19], [51]–[60]. Computed spectra shape and their intense peaks at B3LYP/6-31G(d) level has relatively good agreement with one obtained by experiment (Fig. 6). For instance, maximum peaks in low energy band appeared in 483 nm in comparison with 473 nm in experiment. In following sections effect of other functionals and basis sets for reproduce of UV-Vis. spectrum of DCM have been reported.

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A. *Functional Investigation*

The obtained results for 11 functionals are summarized in Table 2. Obviously, the selected functional impacts significantly on the computed $S_0 \rightarrow S_1$ and $S_0 \rightarrow S_2$ transition energies, that tend to increase when more exact exchange is plugged in an expected trend for π-π* transitions [31], [61]–[63]. As seen in the Table 2, generally in hybrid functionals, with the increase in exact exchange functional ratio (X), a blue shift occurs in electronic transition. For example when X in X3LYP increase about 5% compared with B3LYP, both electronic transition $S_0 \rightarrow S_1$ and $S_0 \rightarrow S_2$ have shown the blue shift but electronic excitation resulted from O3LYP functional that has smaller exact exchange contribution than B3LYP has shown a red shift. Shape of the calculated spectra arising from the use of all 11 functionals by medium size Pople basis 6-31G(d) has been shown in Fig. 8.

Also computed transition energies from B3LYP, X3LYP, mPW3PBE and B97-2 are in agreement with experiment but with slightly underestimation. Although previous studies [64]–[66] about the global hybrids such as PBE0 and B3LYP, have been found efficient within the vertical time dependent DFT approximation for some of dyes but in DCM case the mPW1PBE and mPW1LYP are more efficient. Estimated transition energies by other functionals have large deviation from experimental results. For example computed first electronic transition energies at BLYP, O3LYP and WB97XD levels are in 563, 509 and 411 nm, respectively.

B. *Basis Set Investigation*

The results of the ABS investigation can be found in Table 3 whereas depicted topologies of corresponding spectra placed in Fig. 9. In this examination, a PCM-mPW1PBE method on B3LYP optimized structure has been used. The efficiency of mPW1PBE to discussion of electronic transition in DCM was proved in previous section. One of the vast investigations aiming at assessing the influence of the size of the some Pople basis on the TD-DFT calculation and topology of the vibronic band has been done by Jacquemin [64] and Grimme [63] on the other dyes. At the present work almost all selected basis give similar outcomes analogous to [64]. Range of variation of low energy vertical transition wavelengths ($S_0 \rightarrow S_1$) from compact basis 6-31G to sizable basis 6-31+G(3df,2p) and cc-pVTZ is about 13 nm. As shown in Table 3, effect of diffuse basis function (+) in variation of electronic transition energies is larger than polarization functions (p, d and f basis functions).

Table 2. Computed two lower singlet-singlet excitations of solvated *cis*-DCM by using different functionals. All results obtained at PCM-TDDFT/6-31G(d) level of theory on PCM-B_3LYP/6-31G(d) optimized structure. X and Y are exact exchange and correlation contributions. Osc. is oscillator strength of electronic transition between to states. Transition energy unit reported in nm.

Functional	X	Y	Type	$s_0 \rightarrow s_1$	Osc. $(s_0 \rightarrow s_1)$	$s_0 \rightarrow s_2$	Osc. $(s_0 \rightarrow s_2)$
B3LYP	20	81	HGGA	483.2	1.28	362.4	0.32
X3LYP	21.8	87.1	HGGA	477.7	1.32	357.5	0.31
O3LYP	11.6	78	HGGA	508.8	1.12	387.1	0.37
BLYP	0	100	GGA	562.8	0.84	435.6	0.49
mPW1PBE	25	100	HGGA	468.6	1.38	348.6	0.30
mPW3PBE	20	100	HGGA	483.7	1.29	362.2	0.32
mPW1LYP	25	100	HGGA	468.0	1.37	349.0	0.29
M06-2X	54	100	HMGGA	423.8	1.64	314.8	0.27
B97-2	21	100	HGGA	478.2	1.31	357.6	0.32
WB97XD	22	100	HMGGA	410.83	1.71	308.1	0.26
CAM-B3LYP			HGGA	418.6	1.66	316.9	0.25

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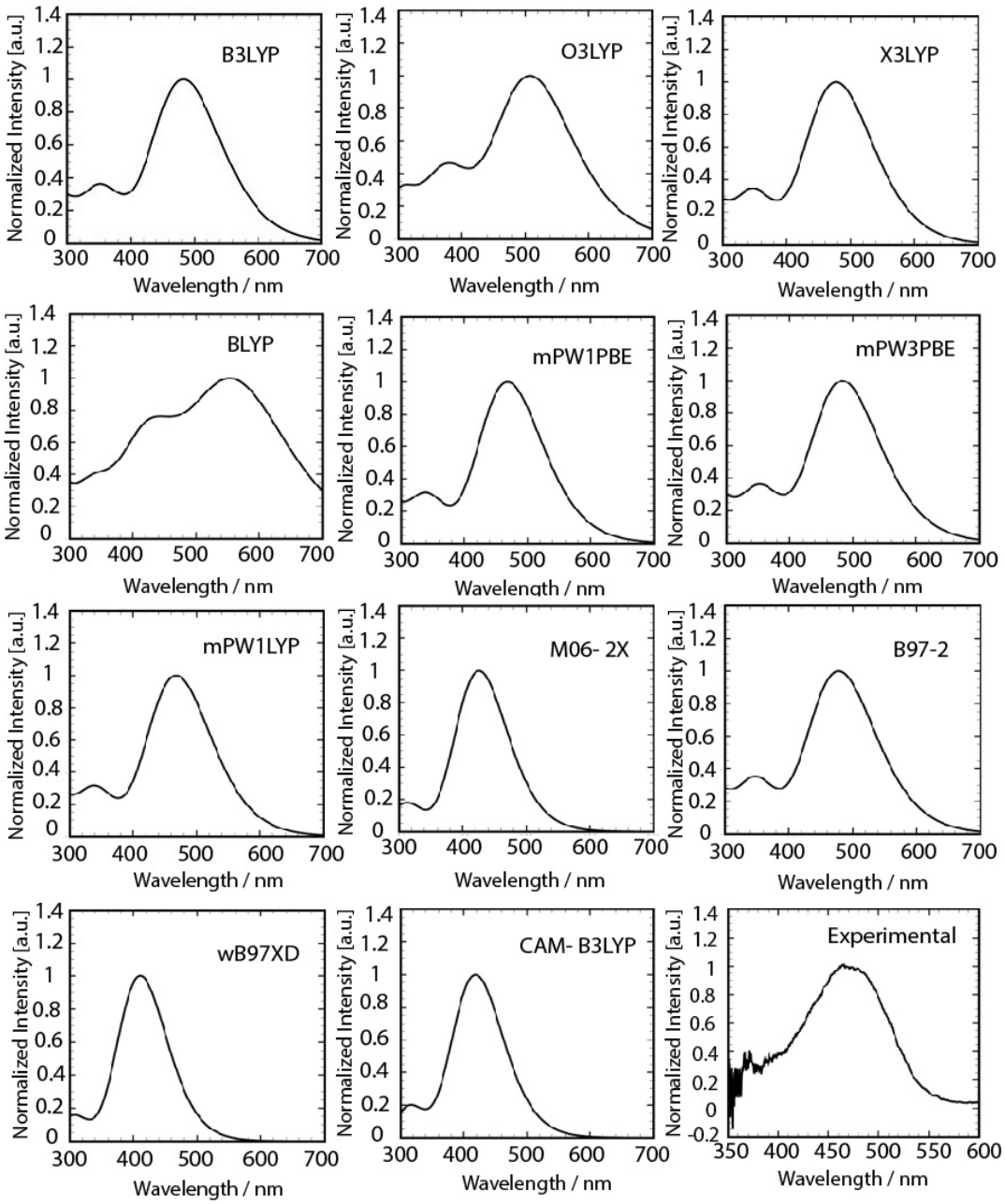

Fig. 8. Calculated absorption spectra for solvated DCM using different functionals and the same basis set along with experimental spectra.

Table 3. Computed two lower singlet-singlet excitation energy of solvated DCM by different basis sets. All results obtained at PCM-TD-mPW1PBE/basis level

Basis set	$s_0 \rightarrow s_1(nm)$	Osc. $(s_0 \rightarrow s_2)$	$s_0 \rightarrow s_2(nm)$	Osc. $(s_0 \rightarrow s_2)$
6-31G	470.3	1.37	345.9	0.35
6-31G(d)	468.6	1.38	348.6	0.30
6-31+G(d)	480.1	1.37	351.0	0.35
6-31++G(d)	481.0	1.37	351.0	0.35
6-31+G(2d,2p)	481.7	1.37	352.2	0.34
6-31+G(3df,2p)	481.6	1.38	353.2	0.33
cc-pVTZ	476.7	1.38	352.2	0.31

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Resulted shapes of spectra by using different basis sets are very similar to each other and experiments (Fig. 9). By comparison of transition intensities which obtained using different basis sets with experimental absorption, it revealed that 6-31G(d) basis set reproduced experimental absorption spectra of DCM and is better than others. As can be seen in Table 3, range of change in computation of high transition energy band ($S_0 \rightarrow S_2$) is about 6.3 nm.

Apart from the first two compact basis sets in Table 3, other chosen basis sets slightly underestimated the transition energy of low energy peak in comparison with experimental absorption. As a compromise between cost and accuracy computations, 6-31G(d) is appropriate. This basis set defines *d* polarization functions on all atoms except for hydrogen atoms. This basis satisfies both correct shape of spectra and transition energies.Computation of Vertical Spectra with Vibrational Couplings

To improve the accuracy of simulations, the vibronic couplings should be evaluated. Vibrational modes that hook up with the electronic transition are not clear in shape of experimental spectra but may change the final position of maximum peaks in spectra. To the best of our knowledge, there are no previous experimental or theoretical investigations aiming at assessing the influence of the vibrational coupling on position of maximum peaks and shape of spectra for DCM dye.

We are going to correct the computed first transition energy by including the vibrational couplings along with electronic transitions. To perform such task, one needs to calculate the Hessian of both the ground state and relevant (generally first) singlet excited state. In practice, this can be done with DFT and TDDFT for ground and excited states respectively. For a reliable calculation, geometry in both ground and excited states were optimized at PCM-mPW1PBE/6-31G(d) and PCM-TD-mPW1PBE/6-31G(d) approach, respectively. The harmonic frequencies of both structures were computed at aforementioned levels. It should be noted that, vertical excitation without including the vibrational couplings, using latter structure appears in 463 nm.

In the interaction of visible light with solvated DCM, it was found a significant elongation of the C-C bond in the ring containing oxygen upon electron excitation. As a consequence, the vibrational frequency of normal mode corresponding to stretch of this bond are obviously displaced to lower frequency (from 1601 (cm^{-1}) in ground state structure to 1547 (cm^{-1}) in excited state structure). Also elongation of bond length was appeared at nitrile bonds, C=N, upon excitation but not noticeably. Finally vibronic absorption spectra of solvated DCM was simulated (Fig. 10) using computed Hessian matrix of minimum structures in initial state, S_0 and final state, S_1 and FC classes code [32], [58], [59]. Fine structures of vibronic spectra were computed although this fine structure is not obvious in the experimental spectrum.

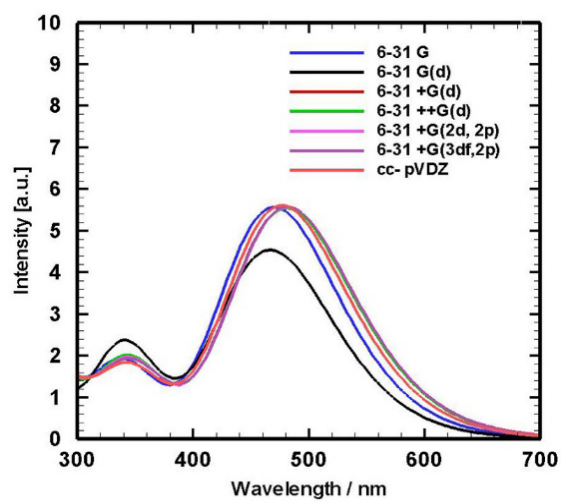

Fig. 9. Computed absorption spectra of solvated DCM using different basis sets and mPW1PBE functional.

To simulating the spectrum line-shape, it is required to convolute the stick-spectrum with a gaussian distribution function with an adequate full-width at the half-maximum (FWHM). Figure 10 shows absorption spectra with different FWHM (FWHM: 1,135 and 2685 (cm^{-1})). By comparing Figure 3 and 10 it

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is appealing that the shape of convoluted spectrum with FWHM equal to 2685 (cm^{-1}) is comparable with experimental one. As can be seen in convoluted spectrum with FWHM equal to 1 (cm^{-1}) the most intense vibronic transition is 0-0 transition in accordance to Frank-Condon rule (i.e. most intense transition occurs from lowest vibrational level, $\nu = 0$, in S_0 electronic state to lowest vibrational level, $\nu = 0$, in S_1 electronic state). The separation between the first and second most intense peaks (in black stick spectrum in Figure 10, FWHM=1 (cm^{-1})) is equal to 101 (cm^{-1}). It can mainly ascribe to third vibration overtone of mode 3 in excited electronic state. Vibrational mode 3 is corresponding to the in planar bending of DCM plane (Fig. 11). Present study predicts the vibronic shape of absorption spectrum for first singlet-singlet electronic transition of DCM in ethanol at 298 (K). Also this is an important result that the computed transition energy was corrected to 472 nm by considering the vibrational couplings during first electronic excitation.

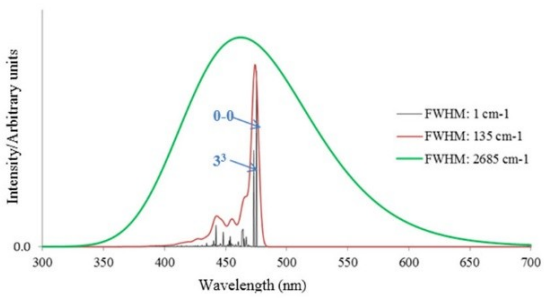

Fig. 10. Computed absorption spectrum of solvated DCM by including the vibrational couplings during electronic excitation. Different full width at half maximum (FWHM) have been chosen in convoluting gaussian (red, green and blue stick line).

Fig. 11. Vibrational mode 3 at first singlet electronic excited state structure of solvated DCM. This mode has been calculated at PCM-TD-mPW1PBE/6-31G(d).

IV. CONCLUSION

The absorption spectra of synthesized DCM dye has obtained using an experimental spectroscopy and a time-dependent density functional theory approach. The room temperature absorption spectrum of DCM exhibit two main bands in polar solvent (ethanol). A low energy band spanned in the 450-520 nm regions and second band with less intensity spanned in higher energy 340 -380 nm. Maximum peak in each aforementioned bands appear in 473 and 362 nm respectively. A theoretical investigation on the electronic structure of DCM is presented in an effort to rationalize our experimental results. It turned out that the selected DFT functional strongly affects the computed transition wavelengths as well as the topology of the absorption spectra. In the present case, mPW1PBE and mPW1LYP provide accurate positions (wavelengths) and intensities for the absorption peaks. These functionals are efficient to determine the shapes of the absorption bands. We evaluated the impact of the atomic basis set on the electronic transition energies using a large panel of Pople's basis sets ,up to the 6-31+G(3df,2p) and also a correlation consistent basis set, cc-pVTZ. On the contrary to functionals, the selected atomic basis set has a relatively modest impact on the photoproperties of DCM. As a compromise between cost and accuracy computations, 6-31G(d) basis set is appropriate for considering the photophysical and photochemical properties of DCM. Thus it was suggested that the mPW1PBE/6-31G(d) gives best results for DCM. This level with low overestimation in computation of transition energy produced both main low and high energy bands with maximum peaks about 463 and 346 nm, respectively. To improve the accuracy of first electronic excitation, vibrational couplings were computed during first electronic transition. It was shown that the most intense vibronic transition is 0-0 transition. Finally the computed transition energy alter from 463 to

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472 nm which was in excellent agreement with experiments.

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